

Ch. 5 Practice Questions

- 1) What is the formula mass of magnesium nitride?
A) 38.32
B) 116.33
C) 90.65
D) 100.95
E) none of the above
- 2) Suppose that in an ionic compound, "M" represents a metal that could form more than one type of ion. In the formula MF_2 , the charge of the M ion would be:
A) -2 B) +1 C) +2 D) +4 E) -1
- 3) What is the name of HIO_3 ?
A) hydroiodic acid
B) iodic acid
C) hydroiodous acid
D) iodous acid
E) none of the above
- 4) What is the proper name for $HBr(aq)$?
A) hydrousbromic acid
B) hydrobromous acid
C) hydrobromic acid
D) bromous acid
E) none of the above
- 5) What is the name of the molecular compound SO_3 ?
A) sulfur trioxide
B) sulfur oxide
C) sulfur tetraoxide
D) sulfur(IV) oxide
E) none of the above
- 6) What is the name of $Ca(NO_3)_2$?
A) calcium nitrite
B) calcium nitrate
C) calcium nitride
D) calcium dinitrite
E) none of the above
- 7) What is the formula for an ionic compound made of magnesium and sulfur?
A) MgS
B) MgS_2
C) Mg_2S_3
D) Mg_2S
E) none of the above

- 8) Which of the following species is a molecular element?
- A) sodium
 - B) neon
 - C) chlorine
 - D) carbon dioxide
 - E) none of the above
- 9) Ammonium fluoride is considered which of the following?
- A) molecular compound
 - B) atomic element
 - C) molecular element
 - D) ionic compound
 - E) none of the above
- 10) How many oxygen atoms are in the formula $\text{Al}_2(\text{CO}_3)_3$?
- A) 3
 - B) 6
 - C) 9
 - D) 1
 - E) none of the above
- 11) What is the oxygen-to-sulfur mass ratio of sulfur dioxide?
- A) 0.5
 - B) 1.0
 - C) 2.0
 - D) 16
 - E) none of the above

Answer Key

Testname: PRACTICEQ_CH05

- 1) D
- 2) C
- 3) B
- 4) C
- 5) A
- 6) B
- 7) A
- 8) C
- 9) D
- 10) C
- 11) B